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Cool! I'am really happy

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#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

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15. A first order reaction takes 20 minutes for 25% decomposition. Calculate the time when 75% of the reaction will be completed. [3]

(Given: $\log 2 = 0.3010$, $\log 3 = 0.4771$, $\log 4 = 0.6021$)

16. It is given: (for 1^{st} order reaction)

$t = 20$ min

$A_0 = 100\%$

$A = 100 - 25 = 75\%$

$k = ?$

$k = \frac{2.303}{t} \log \frac{A_0}{A}$

$\Rightarrow k = \frac{2.303}{20} \log \frac{100}{75}$

$k = \frac{2.303}{20} \log (1.33)$

$\Rightarrow \frac{2.303}{20} = 0.1248 \Rightarrow k = 0.1248 \text{ min}^{-1}$

So for 75% completion of reaction :-

$t = \frac{2.303}{k} \log \frac{A_0}{A}$

$\Rightarrow \frac{1}{0.1248} = \frac{2.303}{20} \times \frac{100}{25} \Rightarrow \frac{1}{0.1248} = 2.303 \times 0.6021 \Rightarrow 96.96 \text{ min.}$

16. The following compounds are given to you:

2-Bromopentane, 3-Bromo-2-methylbutane, 1-Bromopentane

(a) Write the compound which is most reactive towards $\text{S}_{\text{N}}2$ reaction.

(b) Write the compound which is optically active.

(c) Write the compound which is most reactive towards β -elimination reaction. [3]

17. (a) The compound most reactive towards $\text{S}_{\text{N}}2$ reaction :-

$\text{CH}_3\text{-CH}_2\text{-CH}_2\text{-CH}_2\text{-Br}$

(b) The compound which is optically active :-

$\text{CH}_3\text{-CH}_2\text{-CH}(\text{CH}_3)\text{-CH}_2\text{-Br}$

(c) The compound which is most reactive towards β -elimination reaction is

$\text{CH}_3\text{-CH}_2\text{-C}(\text{CH}_3)=\text{CH}_2$

$\text{CH}_3\text{-CH}_2\text{-C}(\text{CH}_3)_2\text{-CH}_2\text{-Br}$

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