

#Jenny



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#Rio



Cool! I'am really happy

#Markus Jensen



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My friends are so mad that they do not know how I have all the high quality ebook which they do not!

#Diego Butler



so many fake sites. this is the first one which worked! Many thanks

Practice Problems

- Calculate the pH and pOH of aqueous solutions having the following ion concentrations.

$$[\text{OH}^-] = 6.5 \times 10^{-6}$$

$$\text{pOH} = -\log[\text{OH}^-] \quad \text{pH} = 14.00 - \text{pOH}$$

$$\text{pOH} = -\log[6.5 \times 10^{-6}] \quad \text{pH} = 14.00 - 5.19$$

$$\text{pOH} = -[\log 6.5 + \log 10^{-6}] \quad \text{pH} = 8.81$$

$$\text{pOH} = -[0.81 + (-6)]$$

$$\text{pOH} = 5.19$$

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